

## LAMPIRAN

### 1. PERHITUNGAN REAGENSIA

1) Pengenceran Larutan Amonia 25% menjadi 2% dalam Etanol 70% sebanyak 500 ml :

$$\text{➤ Volume Amonia} = \frac{500 \text{ ml} \times 2\%}{25\%} = 40 \text{ ml}$$

$$\text{➤ Volume Etanol 70\%} = 500 \text{ ml} - 40 \text{ ml} = 460 \text{ ml}$$

2) Pengenceran Larutan Asam Asetat 99,6% menjadi 10% sebanyak 250 ml :

$$\text{➤ Volume Asam Asetat} = \frac{250 \text{ ml} \times 10\%}{99,6\%} = 25,1 \text{ ml}$$

$$\text{➤ Volume Aquades} = 250 \text{ ml} - 25,1 \text{ ml} = 224,9 \text{ ml}$$

3) Pengenceran Larutan Amonia 25% menjadi 10% dalam Etanol 70% :

$$\text{➤ Volume Amonia} = \frac{100 \text{ ml} \times 10\%}{25\%} = 40 \text{ ml}$$

$$\text{➤ Volume Etanol 70\%} = 100 \text{ ml} - 40 \text{ ml} = 60 \text{ ml}$$

4) Pembuatan Larutan Baku Induk Standar Rhodamin B 1000 ppm :

$$M = \frac{\text{massa}}{\text{volume}}$$

$$1000 \text{ mg/L} = \frac{\text{massa}}{0,1 \text{ L}}$$

$$\text{Massa} = 100 \text{ mg}$$

5) Pembuatan Larutan Baku Antara Standar Rhodamin B 100 ppm :

$$M1 \times V1 = M2 \times V2$$

$$100 \times V1 = 1000 \times 100 \text{ ml}$$

$$V1 = \frac{1000 \times 100 \text{ ml}}{100}$$

V1 = 10 ml (dimasukkan kedalam labu ukur 100 ml ad dengan aqudes sampai tanda batas dan homogenkan)

## 2. PERHITUNGAN ELUEN (N-butanol : Etil asetat : Amonia (10:4:5))

- a) N-butanol :  $\frac{10}{20} \times 100 \text{ ml} = 50 \text{ ml}$
- b) Etil Asetat :  $\frac{4}{20} \times 100 \text{ ml} = 20 \text{ ml}$
- c) Ammonia 10% :  $\frac{5}{20} \times 100 \text{ ml} = 25 \text{ ml}$

## 3. PERHITUNGAN NILAI RF

### ➤ Sampel 1 :

- 1) S1 (+) =  $\frac{4,2}{8} = 0,6$
- 2) S1R1 =  $\frac{4,2}{8} = 0,6$
- 3) S1R2 =  $\frac{4,2}{8} = 0,6$

### ➤ Sampel 2 :

- 1) S2 (+) =  $\frac{4,3}{8} = 0,53$
- 2) S2R1 =  $\frac{4,3}{8} = 0,53$
- 3) S2R2 =  $\frac{4,3}{8} = 0,53$

### ➤ Sampel 3 :

- 1) S3 (+) =  $\frac{4,8}{8} = 0,6$
- 2) S3R1 =  $\frac{3,9}{8} = 0,48$
- 3) S3R2 =  $\frac{3,9}{8} = 0,48$

### ➤ Sampel 4 :

- 1) S4 (+) =  $\frac{4,1}{8} = 0,51$

$$2) S4R1 = \frac{4}{8} = 0,5$$

$$3) S4R2 = \frac{3,9}{8} = 0,48$$

➤ **Sampel 5 :**

$$1) S5 (+) = \frac{3,9}{8} = 0,48$$

$$2) S5R1 = \frac{3,9}{8} = 0,48$$

$$3) S5R2 = \frac{3,9}{8} = 0,48$$

➤ **Sampel 6 :**

$$1) S6 (+) = \frac{5}{8} = 0,62$$

$$2) S6R1 = \frac{5}{8} = 0,62$$

$$3) S6R2 = \frac{5}{8} = 0,62$$

## LAMPIRAN DOKUMENTASI



Organoleptik S1



Organoleptik S2



Organoleptik S3



Organoleptik S4



Organoleptik S5



Organoleptik S6



Sampel 1



Sampel 2



Sampel 3



Sampel 4



Sampel 5



Sampel 6



Pembuatan  
benang wol  
bebas lemak



Proses Maserasi



Waterbath



Eluasi