LAMPIRAN

A. Perhitungan

Diketahui :

- Konsentrasi HCl = 37%
- Berat Molekul HCl = 36, 5 g/mol
- Berat Jenis HCl = 1,19 g/mL
- Konsentrasi Larutan = 5 N
- Valensi HCl = 1
- Volume Larutan = 10 mL

Ditanya : Volume asam klorida yang dibutuhkan berapa mL?

b. Mencari konsentrasi HCl pekat

$$N = \frac{\text{volume larutan x konsentrasi HCl x berat jenis x valensi}}{\text{berat molekul}}$$

$$= \frac{10\ mL \times 37\% \times 1{,}19\ g/mL \times 1}{36{,}5\ g/mol}$$

$$= 12{,}06\ N$$

c. Mencari volume asam klorida pekat yang dibutuhkan

$$N1 \times V1 = N2 \times V2$$

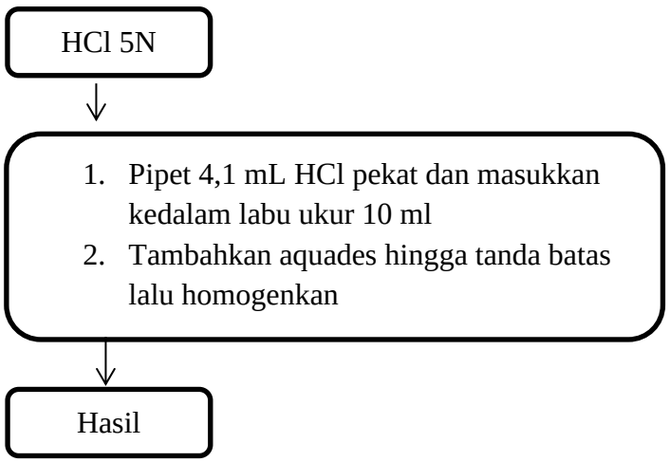
$$12{,}06N \times V1 = 5N \times 10$$

$$V1 = 50 / 12{,}06$$

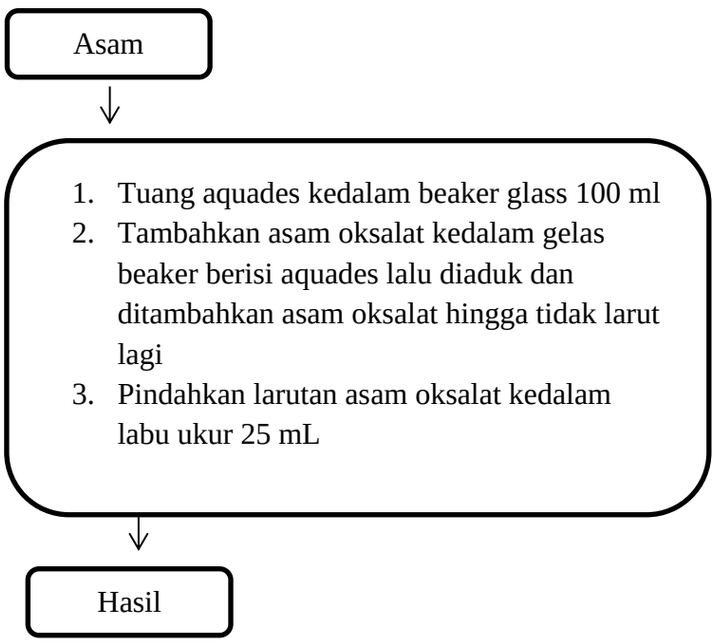
$$= 4{,}1\ mL$$

B. Skema Pembuatan

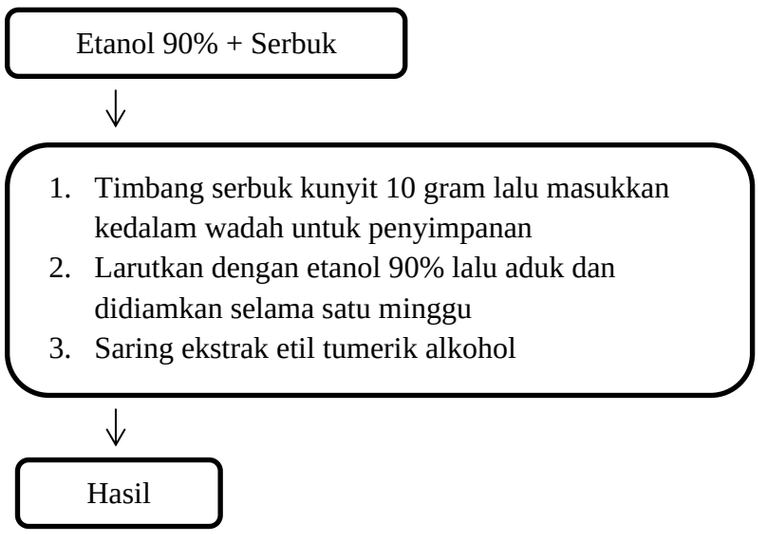
1. Pembuatan larutan HCl 5N



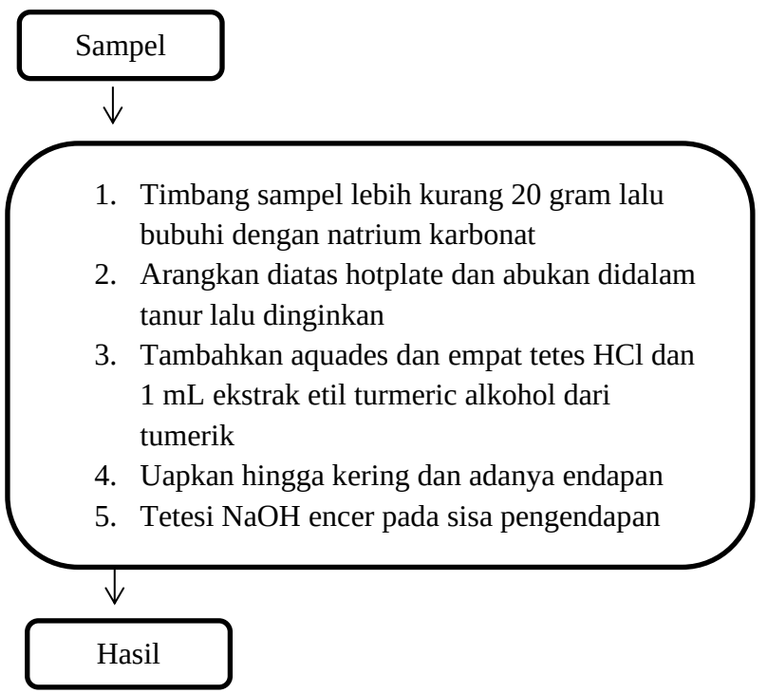
2. Pembuatan Larutan Asam Oksalat Jenuh



3. Pembuatan Ekstrak Etil Alkohol dari Kunyit

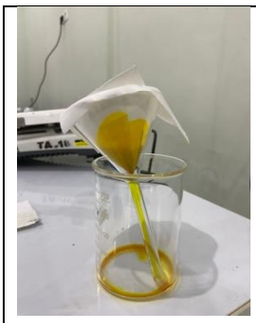


4. Pengujian Boraks secara Kualitatif

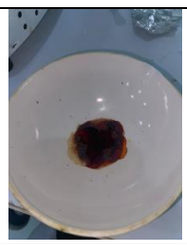
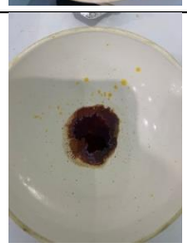


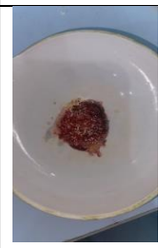
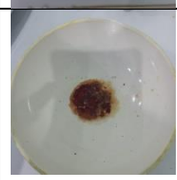
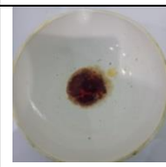
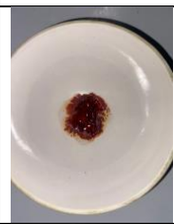
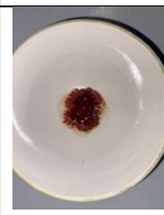
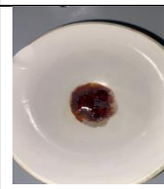
Lampiran 6. 1 Prosedur Kerja

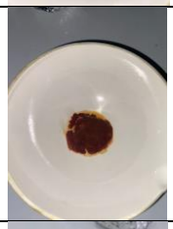
	
Bahan Penelitian	Sampling
	
Sampel + dihabluri Na_2CO_3	Proses Pengarangan Sampel dengan suhu 100°C
	
Proses pengabuan pada Tanur dengan suhu 550°C	Melakukan Proses Desikator atau Pendinginan
	
Larutan Asam Oksalat 10 ml	Larutan Ekstrak Etil Alkohol dari Kunyit

	
Menyaring Larutan Ekstrak Etil Alkohol dari Kunyit	Hasil Pengabuan Sampel

Lampiran 6. 2 Gambar Hasil Pengujian

Sampel	**Setelah Penguapan**	**Hasil Akhir**
KP(Kontrol Positif)		
Pentol 1		
Pentol 1		
Pentol 1		

Cimol 1		
Cimol 1		
Cimol 1		
Cimol 2		
Cimol 2		
Cimol 2		
Pempek 2		

Pempek 2		
Pempek 2		
Cilok 3		
Cilok 3		
Cilok 3		
Sempol 3		
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Sempol 3