LAMPIRAN

Lampiran 1

PERHITUNGAN PEMBUATAN LARUTAN

Perhitungan Standar

Baku Induk I (1000 ppm)

$$Formaldehid\ 1000\ ppm = 1000\frac{\text{mg}}{\text{L}} = \frac{1g}{1000mL} = \frac{0{,}1g}{100mL} = 0{,}1\%$$

V_1 x M_1 = V_2 x M_2

V_1 x 37% = 100 mL x 0,1%

V_1 = 0,27 mL (270 μl)

Jadi volume larutan yang dipipet untuk membuat larutan 1000 ppm sebanyak 0,27 ml.

Baku Induk II (100 ppm)

V_1 x M_1 = V_2 x M_2

V_1 x 1000 ppm = 25 mL x 100 ppm

V_1 = 2,5 mL

Jadi volume larutan yang dipipet untuk membuat larutan 100 ppm sebanyak 2,5 ml.

Baku Kerja = 1 ppm, 1.5 ppm, 2 ppm, 2.5 ppm, 3 ppm (5 $titik$)

Konsentrasi standar

- 1 ppm

  V_1 x M_1 = V_2 x M_2

  V_1 x 100 ppm = 25 mL x 1 ppm

  V_1 = 0,250 mL

  Jadi volume larutan yang dipipet untuk membuat larutan 1 ppm sebanyak 0,250 ml.

- 1,5 ppm

  V1 x M1 = V2 x M2

  V1 x 100 ppm = 25 mL x 1,5 ppm

  V1 = 0,375 mL

  Jadi volume larutan yang dipipet untuk membuat larutan 1,5 ppm sebanyak 0,375 ml.

- 2 ppm

  V1 x M1 = V2 x M2

  V1 x 100 ppm = 25 mL x 2 ppm

  V1 = 0,500 mL

  Jadi volume larutan yang dipipet untuk membuat larutan 2 ppm sebanyak 0,500 ml.

- 2,5 ppm

  V1 x M1 = V2 x M2

  V1 x 100 ppm = 25 mL x 2,5 ppm

  V1 = 0,625 mL

  Jadi volume larutan yang dipipet untuk membuat larutan 2,5 ppm sebanyak 0,625 ml.

- 3 ppm

  V1 x M1 = V2 x M2

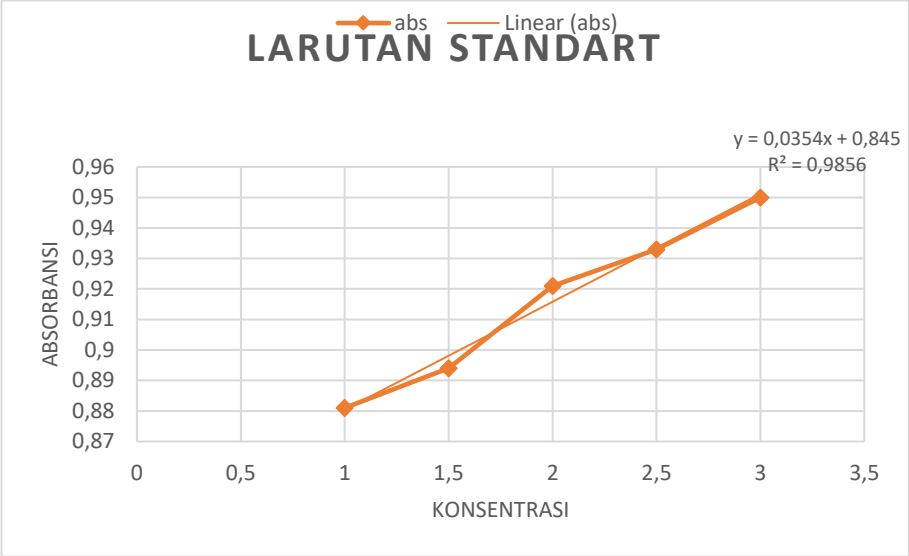
  V1 x 100 ppm = 25 mL x 3 ppm

  V1 = 0,750 mL

  Jadi volume larutan yang dipipet untuk membuat larutan 3 ppm sebanyak 0,750 ml.

Lampiran 2

PERHITUNGAN KADAR



- Kurva Kalibrasi

  $y = ax + b \quad r^2 = 1$

  $y = 0{,}0354x + 0{,}845 \quad r^2 = 0{,}9856$

- Kadar Sampel A1

  $y = 0{,}0354x + 0{,}845$

  $0{,}666 = 0{,}0354x + 0{,}845$

  $0{,}666 - 0{,}845 = 0{,}0354x$

  $\frac{-0{,}179}{0{,}0354} = x$

  $-5{,}0565 = x$

- Kadar Sampel A2

  $y = 0{,}0354x + 0{,}845$

  $0{,}654 = 0{,}0354x + 0{,}845$

  $0{,}654 - 0{,}845 = 0{,}0354x$

  $\frac{-0{,}191}{0{,}0354} = x$

  $-5{,}395 = x$

- Kadar Sampel B1

  $y = 0{,}0354x + 0{,}845$

  $0{,}600 = 0{,}0354x + 0{,}845$

  $0{,}600 - 0{,}845 = 0{,}0354x$

  $\frac{-0{,}245}{0{,}0354} = x$

  $-6{,}920 = x$

- Kadar Sampel B2

  $y = 0{,}0354x + 0{,}845$

  $0{,}640 = 0{,}0354x + 0{,}845$

$$0{,}640\text{-}0{,}845 = 0{,}0354x$$
$$\frac{-0{,}205}{0{,}0354} = x$$
$$-5{,}791 = x$$

- Kadar Sampel C1

$$y = 0{,}0354x + 0{,}845$$
$$0{,}645 = 0{,}0354x + 0{,}845$$
$$0{,}645\text{-}0{,}845 = 0{,}0354x$$
$$\frac{-0{,}200}{0{,}0354} = x$$
$$-5{,}650 = x$$

- Kadar Sampel C2

$$y = 0{,}0354x + 0{,}845$$
$$0{,}604 = 0{,}0354x + 0{,}845$$
$$0{,}604\text{-}0{,}845 = 0{,}0354x$$
$$\frac{-0{,}241}{0{,}0354} = x$$
$$-6{,}808 = x$$

Lampiran 3

Sampel A

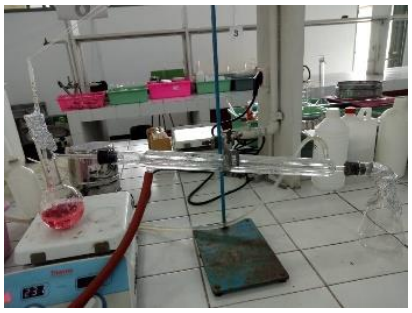
destilasi sampel A

Proses destilasi sampel A

Hasil destilasi sampel A

sampel B

destilasi sampel B

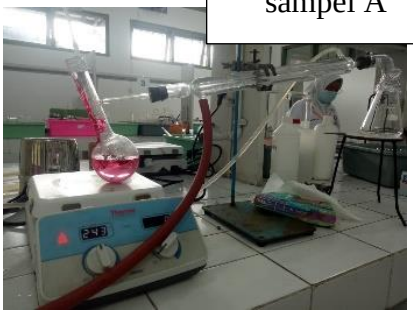


Proses destilasi sampel A

Hasil destilasi sampel A

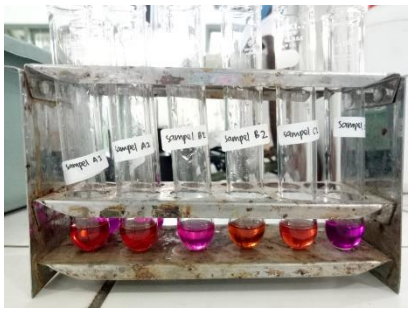
sampel C

destilasi sampel C

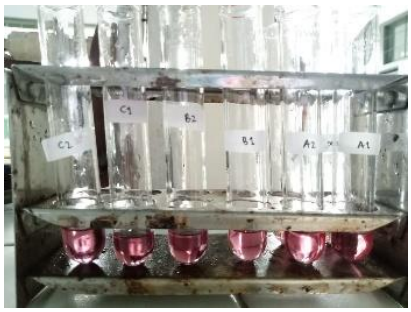


Proses destilasi sampel C

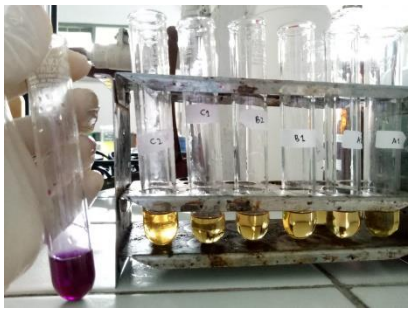
Hasil destilasi sampel A



Uji Kualitatif dengan KMnO4



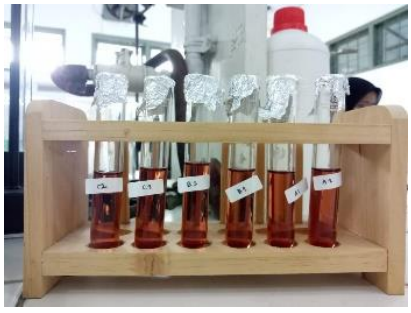
Uji Kualitatif dengan Pereaksi Schiff



Hasil uji sampel dengan pereaksi Schiff dengan HCl

Larutan baku standart

Larutan standart



Pengujian kuantitatif sampel

Uji kuantitatif dengan spektrofotometri uv-vis